



Standard Test Method for Conducting Potentiodynamic Polarization Resistance Measurements¹

This standard is issued under the fixed designation G 59; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ε) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 This test method describes an experimental procedure for polarization resistance measurements which can be used for the calibration of equipment and verification of experimental technique. The test method can provide reproducible corrosion potentials and potentiodynamic polarization resistance measurements.

1.2 *This test method does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 ASTM Standards:

G 3 Practice for Conventions Applicable to Electrochemical Measurements in Corrosion Testing²

G 5 Test Method for Making Potentiostatic and Potentiodynamic Anodic Polarization Measurements²

G 102 Practice for Calculation of Corrosion Rates and Related Information from Electrochemical Measurements²

2.2 Adjunct:

Samples of the Standard AISI Type 430 Stainless Steel (UNS S43000)³

3. Significance and Use

3.1 This test method can be utilized to verify the performance of polarization resistance measurement equipment including reference electrodes, electrochemical cells, potentiostats, scan generators, measuring and recording devices. The test method is also useful for training operators in sample preparation and experimental techniques for polarization resistance measurements.

3.2 Polarization resistance can be related to the rate of general corrosion for metals at or near their corrosion potential, E_{corr} . Polarization resistance measurements are an accurate and

rapid way to measure the general corrosion rate. Real time corrosion monitoring is a common application. The technique can also be used as a way to rank alloys, inhibitors, and so forth in order of resistance to general corrosion.

3.3 In this test method, a small potential scan, $\Delta E(t)$, defined with respect to the corrosion potential ($\Delta E = E - E_{corr}$), is applied to a metal sample. The resultant currents are recorded. The polarization resistance, R_p , of a corroding electrode is defined from Eq 1 as the slope of a potential versus current density plot at $i = 0$ (1-4):⁴

$$R_p = \left(\frac{\partial \Delta E}{\partial i} \right)_{i=0, dE/dt \rightarrow 0} \quad (1)$$

The current density is given by i . The corrosion current density, i_{corr} , is related to the polarization resistance by the Stern-Geary coefficient, B . (3),

$$i_{corr} = 10^6 \frac{B}{R_p} \quad (2)$$

The dimension of R_p is ohm-cm², i_{corr} is $\mu\text{A}/\text{cm}^2$, and B is in V. The Stern-Geary coefficient is related to the anodic, b_a , and cathodic, b_c , Tafel slopes as per Eq 3.

$$B = \frac{b_a b_c}{2.303(b_a + b_c)} \quad (3)$$

The units of the Tafel slopes are V. The corrosion rate, CR , in mm per year can be determined from Eq 4 in which EW is the equivalent weight of the corroding species in grams and ρ is the density of the corroding material in g/cm^3 .

$$CR = 3.27 \times 10^{-3} \frac{i_{corr} EW}{\rho} \quad (4)$$

Refer to Practice G 102 for derivations of the above equations and methods for estimating Tafel slopes.

3.4 The test method may not be appropriate to measure polarization resistance on all materials or in all environments. See 8.2 for a discussion of method biases arising from solution resistance and electrode capacitance.

4. Apparatus

4.1 The apparatus is described in Test Method G 5. It includes a 1 L round bottom flask modified to permit the

¹ This practice is under the jurisdiction of ASTM Committee G01 on Corrosion of Metals, and is the direct responsibility of Subcommittee G 01.11 on Electrochemical Measurements in Corrosion Testing.

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² Annual Book of ASTM Standards, Vol 03.02.

³ Available from ASTM Headquarters. Order PCN 12-700050-00.

⁴ The boldface numbers in parentheses refer to the list of references at the end of this standard.

addition of inert gas, thermometer, and electrodes. This standard cell or an equivalent cell can be used. An equivalent cell must be constructed of inert materials and be able to reproduce the standard curve in Test Method G 5.

4.2 A potentiostat capable of varying potential at a constant scan rate and measuring the current is needed.

4.3 A method of recording the varying potential and resulting current is needed.

5. Test of Electrical Equipment

5.1 Before the polarization resistance measurement is made, the instrument system (potentiostat, X-Y recorder or data acquisition system) must be tested to ensure proper functioning. For this purpose, connect the potentiostat to a test electrical circuit (5). While more complex dummy cells are sometimes needed in electrochemical studies, the simple resistor shown in Fig. 1 is adequate for the present application.

5.2 Use $R = 10.0 \Omega$. Set the applied potential on the potentiostat to $E = -30.0 \text{ mV}$ and apply the potential. The current should be 3.0 mA by Ohm's Law, $I = E/R$.

NOTE 1—When polarization resistance values are measured for systems with different corrosion currents, the value of R should be chosen to cover the current range of the actual polarization resistance measurement. Expected corrosion currents in the microampere range require $R = 1$ to $10 \text{ k}\Omega$.

5.3 Record the potentiodynamic polarization curve at a scan rate of 0.6 V/h from $\Delta E = -30 \text{ mV}$ to $\Delta E = +30 \text{ mV}$ and back to $\Delta E = -30 \text{ mV}$. The plot should be linear, go through the origin, and have a slope 10Ω . The curves recorded for the forward and reverse scans should be identical.

5.4 If the observed results are different than expected, the electrochemical equipment may require calibration or servicing in accordance with the manufacturer's guidelines.

6. Experimental Procedure

6.1 The $1.0 \text{ N H}_2\text{SO}_4$ test solution should be prepared from American Chemical Society reagent grade acid and distilled water as described in Test Method G 5. The standard test cell requires 900 mL of test solution. The temperature must be maintained at 30°C within 1° .

6.2 The test cell is purged at $150 \text{ cm}^3/\text{min}$ with an oxygen-free gas such as hydrogen, nitrogen, or argon. The purge is started at least 30 min before specimen immersion. The purge continues throughout the test.

6.3 The working electrode should be prepared as detailed in Test Method G 5. The experiment must commence within 1 h of preparing the electrode. Preparation includes sequential wet polishing with 240 grit and 600 grit SiC paper. Determine the

surface area of the specimen to the nearest 0.01 cm^2 and subtract for the area under the gasket (typically 0.20 to 0.25 cm^2).

6.4 Immediately prior to immersion the specimen is degreased with a solvent such as acetone and rinsed with distilled water. The time delay between rinsing and immersion should be minimal.

NOTE 2—Samples of the standard AISI Type 430 stainless steel (UNS S45000) used in this test method are available to those wishing to evaluate their equipment and test procedure from Metal Samples, P.O. Box 8, Mumford, AL 36268.

6.5 Transfer the test specimen to the test cell and position the Luggin probe tip 2 to 3 mm from the test electrode surface. The tip diameter must be no greater than 1 mm .

6.6 Record the corrosion potential E_{corr} after 5 and 55-min immersion.

6.7 Apply a potential 30 mV more negative than the recorded 55 min corrosion potential (See Note 3).

NOTE 3—Practice G 3 provides a definition of sign convention for potential and current.

6.8 One minute after application of the -30 mV potential, begin the anodic potential scan at a sweep rate of 0.6 V/h (within 5%). Record the potential and current continuously. Terminate the sweep at a potential 30 mV more positive than the 55 min corrosion potential.

6.9 Plot the polarization curve as a linear potential-current density plot as shown in Practice G 3. Determine the polarization resistance, R_p , as the tangent of the curve at $i=0$.

7. Report

7.1 Report the following information:

7.1.1 The 5 and 55 min corrosion potentials and the polarization resistance value,

7.1.2 Duplicate runs may be averaged, and

7.1.3 Note any deviation from the procedure or test conditions established in this test method.

8. Precision and Bias

8.1 *Precision*—Precision in this test method refers to the closeness of agreement between randomly selected measured values. There are two aspects of precision, repeatability and reproducibility. Repeatability refers to the closeness of agreement between measurements by the same laboratory on identical Type 430 stainless steel specimens repeated with as close as possible adherence to the same procedure. Reproducibility refers to the closeness of agreement between different laboratories using identical Type 430 stainless steel specimens and

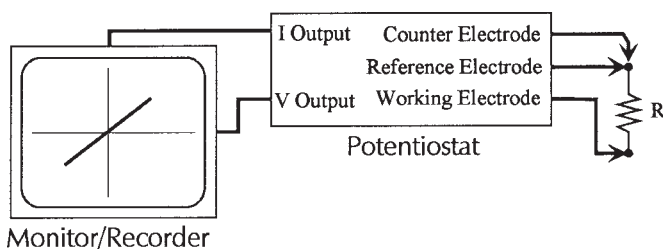


FIG. 1 Arrangement for Testing of Electrical Equipment (Potentiostat, X-Y Recorder)

the procedure specified. An interlaboratory test program with 13 laboratories participating and two, three or four replicate measurements was carried out to establish the precision. The measured values included (Table 1) the corrosion potential measured after 5 and 55 min and the polarization resistance. A research report has been filed with the results of this program.

8.1.1 *Repeatability*— The lack of repeatability is measured by the repeatability standard deviation s_r . The 95 % confidence interval was calculated as $\pm 2.8 s_r$. The values obtained are shown in Table 2. The 95 % confidence interval refers to the interval around the average that 95 % of the values should be found.

TABLE 1 Interlaboratory Test Program Polarization Data for Stainless Steel Type 430 in 1.0 N H₂SO₄ at 30°C

Laboratory	E_{corr} -5min (mV)	E_{corr} -55min (mV)	R_p (ohm-cm ²)
1	-0.519	-0.506	6.47
	-0.519	-0.505	5.88
2	-0.542	-0.521	5.95
	-0.540	-0.519	5.04
3	-0.524	-0.513	6.93
	-0.520	-0.508	6.40
4	-0.555	-0.545	7.70
	-0.565	-0.545	7.70
5	-0.539	-0.524	7.58
	-0.530	-0.510	6.18
6	-0.519	-0.510	7.60
	-0.522	-0.512	7.16
	-0.521	-0.509	6.65
7	-0.522	-0.510	9.06
	-0.520	-0.511	7.07
	-0.523	-0.510	5.85
8	-0.520	-0.508	7.11
	-0.520	-0.508	7.52
	-0.521	-0.510	6.94
9	-0.529	-0.513	7.11
	-0.530	-0.513	7.22
	-0.529	-0.514	7.19
	-0.529	-0.515	7.19
10	-0.514	-0.505	5.17
	-0.516	-0.506	6.90
11	-0.543	-0.529	5.07
	-0.538	-0.524	4.64
12	-0.520	-0.505	5.63
	-0.519	-0.507	6.16
13	-0.531	-0.519	5.08
	-0.529	-0.517	5.38
	-0.529	-0.517	5.90

TABLE 2 Repeatability Statistics

	Average	S_r	95 % Confidence Interval
E_{corr} , 5 min, mV versus SCE	-0.5287	0.00260	± 0.0073 V
E_{corr} , 55 min, mV versus SCE	-0.5151	0.00273	± 0.0076 V
R_p , ohm-cm ²	6.46	0.713	± 2.00 ohm-cm ²

8.1.2 *Reproducibility*— The lack of reproducibility is measured by the reproducibility standard deviation, s_R . The 95 % confidence interval was calculated as $\pm 2.8 s_R$. The values obtained are shown in Table 3.

TABLE 3 Reproducibility Statistics

	Average	S_R	95 % Confidence Interval
E_{corr} , 5 min, mV versus SCE	-0.5287	0.0127	± 0.0356 mV
E_{corr} , 55 min, mV versus SCE	-0.5151	0.0111	± 0.0311 mV
R_p , ohm-cm ²	6.46	1.01	± 2.83 ohm-cm ²


8.2 *Bias*—The polarization resistance as measured by the Test Method G 59 has two sources of bias. The potentiodynamic method includes a double layer capacitance charging effect that may cause the polarization resistance to be underestimated. There is also a solution resistance effect that may cause the polarization resistance to be overestimated. This bias will depend on the placement of the reference electrode and electrolyte conductivity. Refer to Practice G 102 for further discussion on the effects of double layer capacitance and solution resistance on polarization resistance measurements.

9. Keywords

9.1 anodic polarization; auxiliary electrode; cathodic polarization; corrosion; corrosion potential; corrosion rate; current density; electrochemical cell; electrochemical potential; Luggin probe; mixed potential; open-circuit potential; overvoltage; polarization resistance; potentiodynamic; reference electrode; solution resistance; Stern-Geary coefficient; Tafel slope; working electrode

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